

Day 25 - Good Morning!

ionic compounds \Rightarrow salts

- periodic table

\Rightarrow elements in the same column have similar properties

\Rightarrow noble gases - not reactive.

\Rightarrow stable number of electrons

\Rightarrow All elements try to get the same number of electrons as a noble gas.

	element	# prot.	# elec
	F	9	9
$\text{LiO} \Rightarrow \text{Li}_2\text{O}$ +1 -2	F ⁻	9	10
	Cl	17	17
$\text{KF} \Rightarrow \text{K}_2\text{F}$ +1 -1	Cl ⁻	17	18
	Na	11	11
$\text{CaCl} \Rightarrow \text{CaCl}_2$ +2 -1	Na ⁺	11	10
	Mg	12	12
$\text{CsI} \Rightarrow \text{CsI}$ +1 -1	Mg ²⁺	12	10
	O	8	8
	O ²⁻	8	10

1st col. \Rightarrow +1

2nd col. \Rightarrow +2

halogens \Rightarrow -1

②

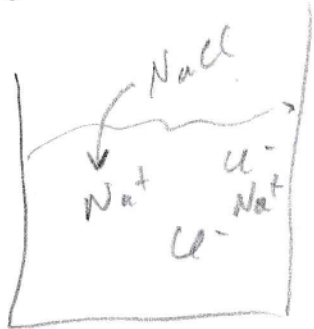
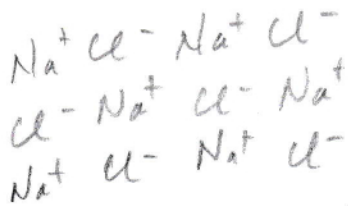


Na transfers an elec. to Cl.

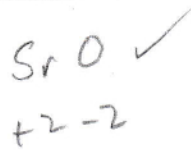
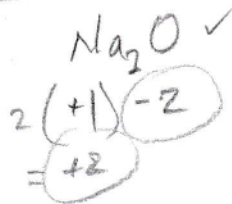
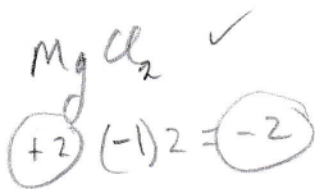
⇒ creates ions

⇒ ionic compound

- stay together because opposite charges attract.



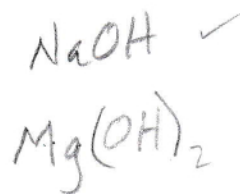
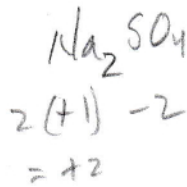
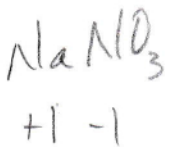
⇒ no molecules



polyatomic ions.



OH^-
↑ hydroxide



- LiO
- KF
- CaCl
- NaBr

