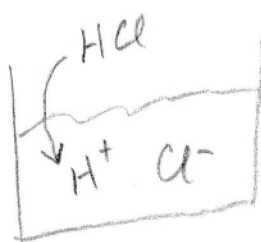
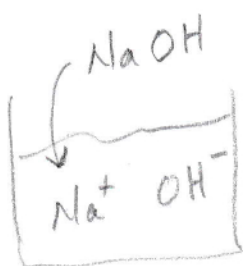
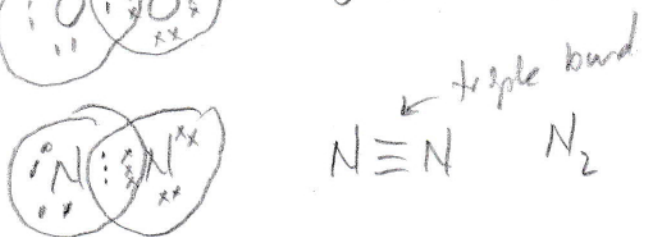
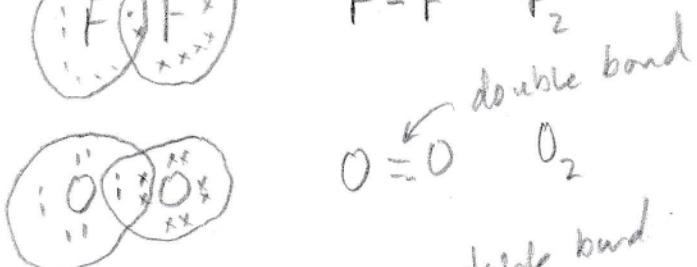
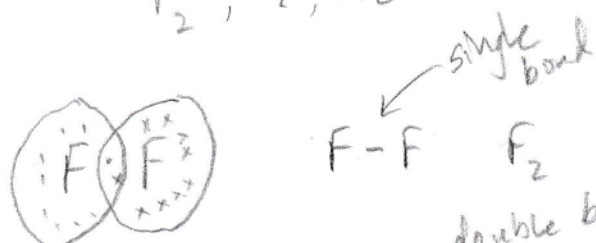
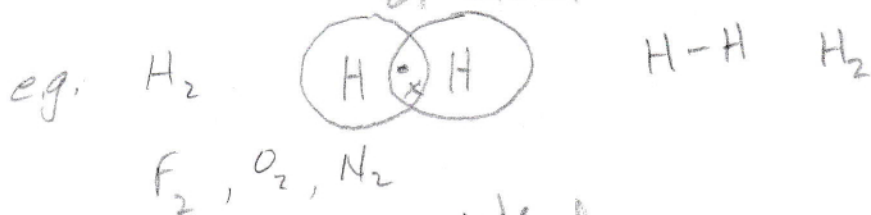


Day 26 - Good Morning!

ionic bond \Rightarrow transfer electrons
 \Rightarrow opposite charges attract

covalent bond \Rightarrow share electrons

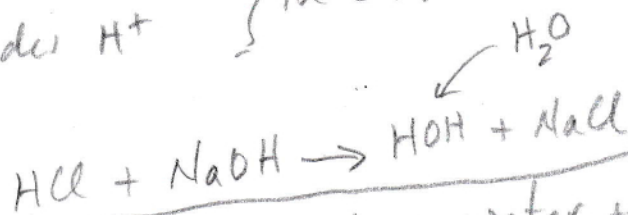
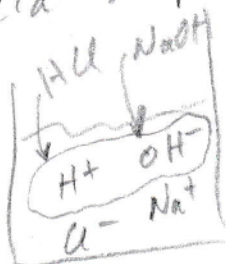
(to get the same number of elec. as a noble gas)



When acids are placed in water, they act like ionic compounds and separate into ions.
 - always an H^+



base \Rightarrow provides OH^-
 acid \Rightarrow provides H^+ } in water



acid + base gives water + salt

③

acid concentration (molarity)

$\text{conc.} \Rightarrow 10^{-1} \quad 10^{-2} \quad 10^{-3} \dots 10^{-7} \dots 10^{-12} \quad 10^{-13} \quad 10^{-14}$
 more acidic ↑ neutral more basic
 num # $\rightarrow 1 \quad 2 \quad 3 \dots 7 \dots 12 \quad 13 \quad 14$
 (pH)

(normal conditions)

$$10^{-1} = \frac{1}{10} = .1$$

$$10^{-2} = \frac{1}{10^2} = \frac{1}{100} = .01$$

$$10^{-3} = \frac{1}{10^3} = \frac{1}{1000} = .001$$

exponent
 $\Rightarrow \log$

$$\text{pH} = 5 \Rightarrow 10^{-5} \Rightarrow \text{acidic}$$

$$\text{pH} = 3.3 \Rightarrow 10^{-3.3} \Rightarrow \text{acidic}$$

$$\text{conc.} = 10^{-9} \Rightarrow 9 \Rightarrow \text{basic}$$

$$\text{pH} = 8 \Rightarrow 10^{-8} \Rightarrow \text{basic}$$

acid-base indicator

\Rightarrow change colors, one color in acid
different color in base

